

Math Diversion Problem 897

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Zathras is used to being beast of burden to other
people's needs. Very sad life... Probably have
very sad death. But, at least
there is symmetry.

— Zathras (a character on Babylon 5)

Source: <https://www.alvinisd.net/cms/lib03/TX01001897/Centricity/Domain/4240/practice%20test%20stoich.pdf>

Title: Mystery Hydrocarbon C_xH_y, Prob. 14, p.2

Presenter: Patrick

Definitions:

FW = Formula weight = molar mass

ppt = precipitate

At wt = atomic weight

1 Problem

A given sample of some hydrocarbon is burned completely and it produces 0.44 g of CO₂ and 0.27 g of H₂O. Determine the empirical formula of the compound.

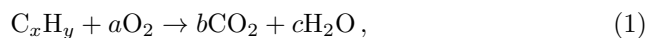
- a. CH b. C₂H₃ c. CH₂
d. C₂H₅ e. CH₃

Ans: e.

2 SOLUTION

Step 1.

We'll begin by finding a balanced equation, starting with



where a, b, c are positive integer coefficients to be solved for in terms of unknowns x and y . To that end, we must have the coefficients cooperate with balancing the carbon, hydrogen, and oxygen:

$$\begin{aligned} \text{C} &: x = b \\ \text{H} &: y = 2c \\ \text{O} &: 2a = 2b + c \end{aligned} \tag{2}$$

On solving a, b, c in terms of x and y , we get

$$\begin{aligned} a &= x + y/4 \\ b &= x \\ c &= y/2 \end{aligned} \tag{3}$$

Step 2. Once again, a diagram.

	Molar Mass (g/mol):		44.01	18.01
Elements/ Compounds:	<div style="border: 1px solid black; display: inline-block; padding: 2px 10px;">C_xH_y</div> + <div style="border: 1px solid black; display: inline-block; padding: 2px 10px;">O_2</div>	→	<div style="border: 1px solid black; display: inline-block; padding: 2px 10px;">CO_2</div> + <div style="border: 1px solid black; display: inline-block; padding: 2px 10px;">H_2O</div>	
MoleStats:	1	$x+y/4$	x	$y/2$
Mass (g):			0.44	0.27
Moles:			┌-0.01	┌-0.015

Figure 1. This sparse diagram is enough to find the ratio of x to y .

Step 3.

Next, we write down our mole proportion between columns 3 and 4:

$$\frac{x}{y/2} = \frac{0.01}{0.015} \tag{4}$$

But this is only one equation for two variables. No, we don't need another equation because all we need is the ratio of the two variables this time. Solving for x/y , we get a 1/3. Therefore, the empirical formula is CH_3 .

3 Appendix

There are four main types of data in the stoich diagrams I make. The most common are data from given information, from the coefficients of the balanced equation, and from data tables, such as a periodic table of elements for molar mass information. This kind of data I do not mark up. The second kind of data in stoich diagrams comes from computations based on data in the same column, for which I use the turnstile (┌) to indicate. The third kind of data is a result in one column that required data from at least one other column to calculate it, and I indicate that kind of value or result by use of the underlining. The fourth kind of data in the figures is the result of combining given information to derive a secondary value. I indicate this kind of data with a right arrowhead (►).